

Chapter 11 Review Molecular Composition Of Gases Section 11 2 Answers

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CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L.

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Chapter 11 - Molecular Composition of Gases 11-1 Volume-Mass Relationships of Gases I. Measuring and Comparing the Volumes of Reacting Gases A. Observations of Gay-Lussac 1. 2 liters H₂ + 1 liter O₂ → 2 liters H₂O vapor 2. 2 volumes H₂ + 1 volume O₂ → 2 volumes H₂O vapor 3. 1 volume H₂ + 1 volume Cl₂ → 2 volumes HCl 4.

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Molar Conversions & Stoichiometry. SOL Review #5 Chapters 10, 11 Name: ____ Chapter 10: Molar Conversions Molar Mass Mole, Gram, Liter, Particles conversions Empirical/Molecular Formulas. Percent Composition Chapter 11: Stoichiometry Molar Ratios Mole-Mole **very common! Gram-Gram L-L

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CHAPTER 11 REVIEW Molecular Composition of Gases MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. The average speed of a gas molecule is most directly related to the . (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) number of moles in the sample 2.

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334 CHAPTER 11 FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas they contain. Hydrogen molecule 1 mol H₂ at STP = 22.4 L Oxygen molecule 1 mol O₂ at STP = 22.4 L Carbon dioxide molecule 1 mol CO₂ at STP = 22.4 L

L-L

Given the number of grams of each atom, find the empirical formula. Given additionally the molar mass, determine the molecular formula. Determine the percentage composition for a substance given the total and element masses. Chemistry Chapter 11 Test Review Define stoichiometry. What does stoichiometry do for you? How is the molar ratio determined?

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CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances; ____ a. temperature increases, volume stays the same

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