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Chapter 12 - Stoichiometry

Chapter 11 Small-Scale Lab Section 11.3 Precipitation Reactions: Formation of Solids, page 345 Analysis 1. a. $\text{Na}_2\text{CO}_3 + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{CO}_3(\text{s})$ b. $2\text{Na}_3\text{PO}_4 + 3\text{Pb}(\text{NO}_3)_2 \rightarrow 6\text{NaNO}_3 + \text{Pb}_3(\text{PO}_4)_2$

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Pb 3(PO 4) 2(s) 2.

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e. 12 g 7. How many grams of nitric acid, HNO_3 , can be prepared from the reaction of 138 g of NO_2 with 54.0 g H_2O according to the equation below? $3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO}$...
Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers: $2\text{NH}_3 + \frac{7}{2}\text{O}_2 \rightarrow 2\text{NO}_2 + 3\text{H}_2\text{O}$. The ...

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