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# Chapter 14 Review Acids And Bases Mixed

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1. Name the following compounds as acids: sulfuric acid a.  $\text{H}_2\text{SO}_4$  sulfurous acid b.  $\text{H}_2\text{SO}_3$  hydrosulfuric acid c.  $\text{H}_2\text{S}$  perchloric acid d.  $\text{HClO}_4$  hydrocyanic acid e. hydrogen cyanide
2.  $\text{H}_2\text{S}$  Which (if any) of the acids mentioned in ...

## Chapter 14: Acids and Bases

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water. stage 1:  $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$  b. Which stage of ionization usually produces more ions? Explain your answer.

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Interactive Example 14.1 - Acid Dissociation (Ionization)  
Reactions Write the simple dissociation (ionization) reaction (omitting water) for each of the following acids: a. Hydrochloric acid (HCl) b. Acetic acid ( $\text{HC}_2\text{H}_3\text{O}_2$ ) c. The anilinium ion ( $\text{C}_6\text{H}_5\text{NH}_3^+$ ) d. The hydrated aluminum(III) ion  $[\text{Al}(\text{H}_2\text{O})_6]^{3+}$

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Arrhenius Acid -a hydrogen containing compound that ionizes to produce hydrogen ions (  $H^+$  ) when dissolved in water Because molecular acids are not made of ions, they cannot dissociate.  $HCl(aq) \rightarrow H^+(aq) + Cl^-(aq)$  They must be pulled apart, or ionized, by the water. Chapter 14: Acids and Bases Ch 14 Page 1

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ashleysanders47. Terms in this set (51) Acids. reacts with base to form salt and water. Starts with  $\text{H}^+$  (hydrogen ion). Binary Acid. an acid that contains only two different elements: hydrogen and one of the more electronegative ...

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