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CHAPTER 10: Chemical Quantities

BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named

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Avagadro, and he decided to use the

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number.  $6.02 \times 10^{23}$  representative particles of a substance, named in honor of Italian scientist Amedeo Avogadro di Quaregna.

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reactants and products.

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## *CHAPTER 10: Chemical Quantities*

Chapter 10 - Chemical Quantities 1.

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Chemical Quantities Or How I Learned to Love the Mole 3. The Mole How do you measure matter? If it is something large like the number of beans in a bag, you just count them. If, however, it is something exceedingly small like an atom or ion, how do you count it?

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(in 1 formula unit, there are 12 oxygen atoms (4x3), so in 10 formula units there are 120 oxygen atoms (10x12))

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Chapter 10 "Chemical Quantities" Yes ,  
you will need a calculator for this

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10. Look at Figure 10.16 and Table 10.3. Name three compounds that have an empirical formula of CH. 11. Fill in the labels on the diagram below.

MICROSCOPIC INTERPRETATION

MACROSCOPIC INTERPRETATION  $\text{SO}_3$   
molecule 1 mol  $\text{SO}_3$   $\text{SO}_3$  composed of  
composed of S atom and sulfur atoms (1023) oxygen atoms 3 atoms and

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CHAPTER 10, Chemical  
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