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The formula I determined in my experiment is an empirical formula and not a molecular formula because if it was a molecular formula I would have left is as $\text{Cu}_{0.0065950}$. with no dividing by the lowest ratio and rounding it up to a whole number.

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The sequence of 35 experiments follows

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topics covered in most general chemistry texts as outlined in a correlation chart to several of the publisher's general chemistry titles. The fifth edition adds the synthesis and empirical formula determination of zinc iodide, and deletes the use of iron.

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The sequence of experiments has been refined to follow topics covered in most general chemistry textbooks. In addition, Murov has included a correlation chart that links the experiments in the manual

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to the corresponding chapter topics in several Cengage Learning general chemistry titles.

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second semester of a two semester series for chemistry majors and others requiring experimental physical chemistry laboratory experience. The course is largely structured around the standard experimental physical chemistry text by Garland, Nibler and Shoemaker (Experiments in Physical Chemistry, 6th Ed., McGraw-Hill, Boston, 2003). Detailed ...

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