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conductor of electricity, but the addition of an acid or base produces ... distinguish solutions of acids and bases, as shown in Figure 18.2. Aqueous solutions of acids cause blue litmus paper to turn red. Aqueous

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strength of a weak acid with the strength of its conjugate base and the strength of a weak base with the strength of its conjugate acid.

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acids and bases. A Lewis acid is an electron pair acceptor and a Lewis base is an electron pair donor. A Lewis acid may not have an ionizable hydrogen ion or hydroxide ion to qualify as an Arrhenius acid or base. A Lewis acid may not have a hydrogen ion to donate, so that it could not qualify as a Bronsted-Lowry acid. However, all Lewis bases are

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Study Guide: Acids and Bases 1) List at least three characteristic properties of acids and three of bases. ACIDS pH less than 7 Have a sour taste Change the color of many indicators Are corrosive (react with metals) Neutralize bases Conduct an electric current BASES

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- Identify the physical and chemical properties of acids and

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bases. • Classify solutions as acidic, basic, or neutrga. • Compare the Arrhenius and Brønsted-Lowry models of acids and bases. Lesson Resources Section Focus Transparency 71 and Master Teaching Transparency 55 and Master Study Guide for Content Mastery, p. 109 TCR Multimedia ...

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of acid or base are added. A buffer system should contain considerably more acid than base. Specific buffer systems should be chosen based on the pH that must be maintained. 114
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