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Example Problem Applying the Henderson-Hasselbalch Equation Calculate the pH of a buffer solution made from 0.20 M $\text{HC}_2\text{H}_3\text{O}_2$ and 0.50 M $\text{C}_2\text{H}_3\text{O}_2^-$ - that has an acid dissociation constant for $\text{HC}_2\text{H}_3\text{O}_2$ of 1.8×10^{-5} .

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Microeconomic Theory: Henderson, James M., Quandt, Richard ...

In chemistry and biochemistry, the Henderson-Hasselbalch equation $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ can be used to estimate the pH of a buffer solution. The numerical value of the acid dissociation constant, K_a , of the acid is known or assumed. The pH is calculated for given values of the concentrations of the acid, HA and of a salt, MA, of its conjugate base, A^- ; for example, the solution may contain acetic ...

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